TFY4115 Fysikk

Faseoverganger (smelte, fordampe)

Y&F kap.17.6+18.6 (8 sider)
L&H&L Kap. 17.10 (1½ side)
H&S kap 10 (6 sider)

Varme Q tilført et legeme kan:

1) Varme opp stoff: $Q = C \cdot n \cdot \Delta T$

der C = molar varmekapasitet

2) Smelte stoff: $Q = L'_s \cdot \Delta m$

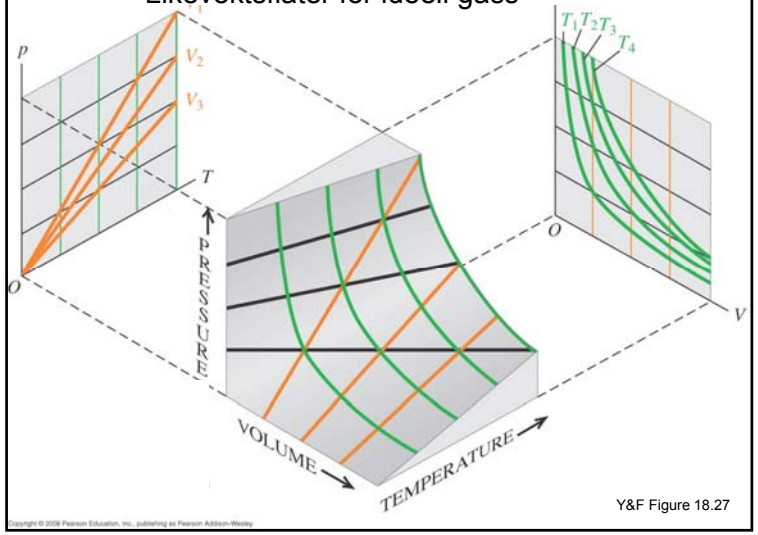
der L'_s = spesifikk smeltevarme (J/kg)

3) Fordampe stoff: $Q = L'_f \cdot \Delta m$

der L'_f = spesifikk fordampingsvarme (J/kg)

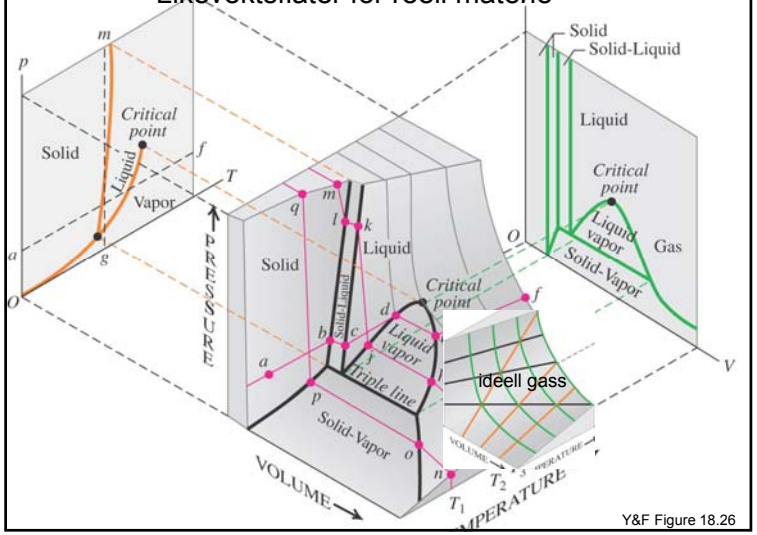
4) Utvide en gass isotermt $Q = \int p \, dV$

Likevektsflater for ideell gass

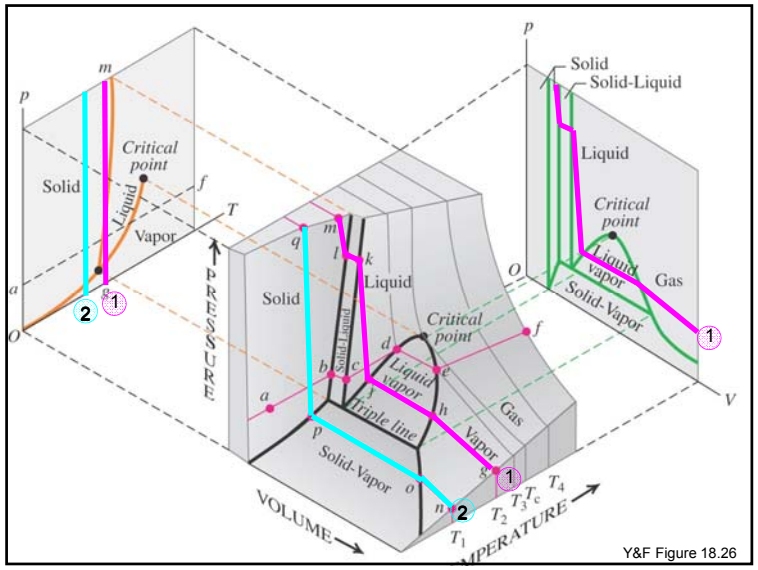


Y&F Figure 18.27

Likevektsflater for reell materie

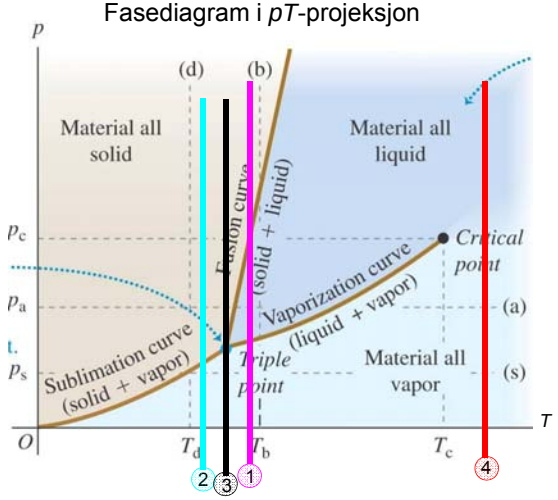


Y&F Figure 18.26



Y&F Figure 18.26

Fasediagram i *pT*-projeksjon



Y&F Figure 18.24

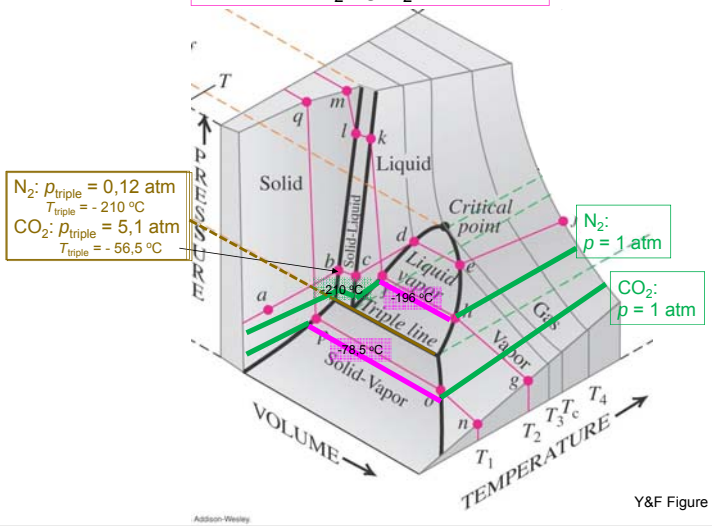
Table 18.3 Triple-Point Data

Substance	Temperature (K)	Pressure (Pa)
Hydrogen	13.80	0.0704×10^5
Deuterium	18.63	0.171×10^5
Neon	24.56	0.432×10^5
Nitrogen	63.18 = - 210 °C	0.125×10^5 = 0,12 atm
Oxygen	54.36	0.00152×10^5
Ammonia	195.40	0.0607×10^5
Carbon dioxide	216.55 = - 56,5 °C	5.17×10^5 = 5,1 atm *)
Sulfur dioxide	197.68	0.00167×10^5
Water	273.16 = 0,01 °C	0.00610×10^5 = 0,0060 atm

*) ved 1 atm:
sublimerer ved -78,5 °C

Y&F Table 18.3

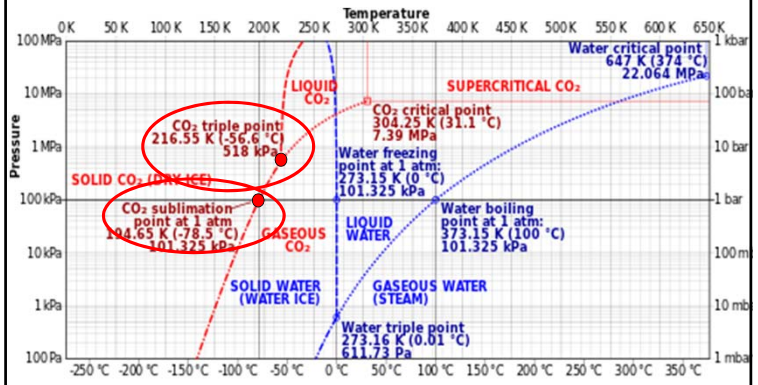
Likevekt CO_2 og N_2 ved 1 atm



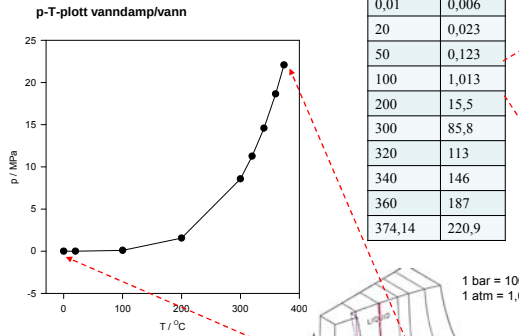
Y&F Figure 18

pT-fasediagram for CO_2 og H_2O

Fra: http://en.wikipedia.org/wiki/Dry_ice



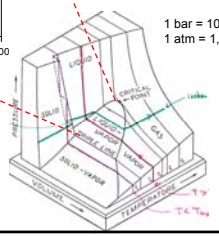
Vanndamptrykk p sfa. temp. T (kokepunkt)



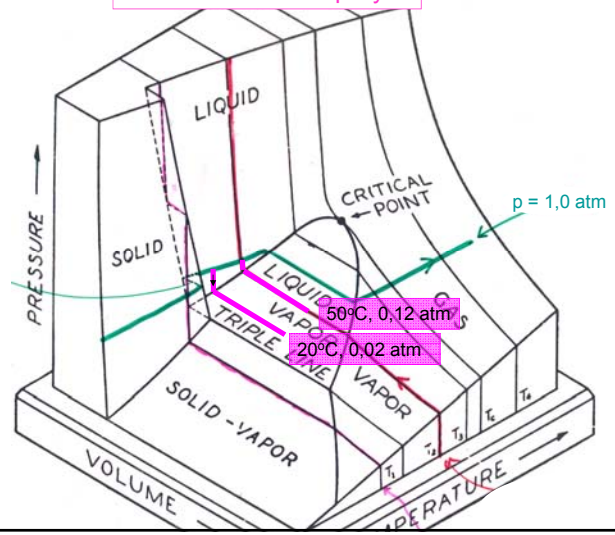
T / °C	p / bar
0,01	0,006
20	0,023
50	0,123
100	1,013
200	15,5
300	85,8
320	113
340	146
360	187
374,14	220,9

T / °C	p / bar
90	0,701
95	0,846
100	1,013
105	1,22
110	1,43

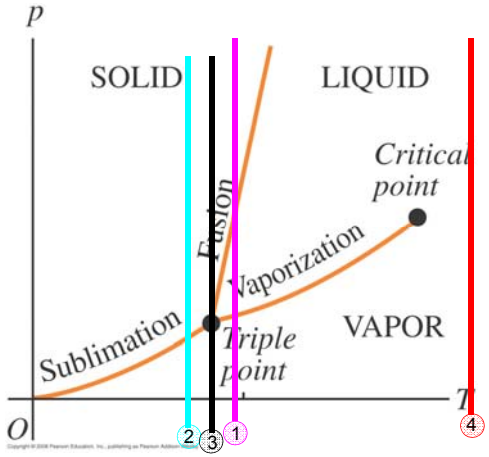
1 bar = 100 kPa = 0,1 MPa
1 atm = 1,013 bar

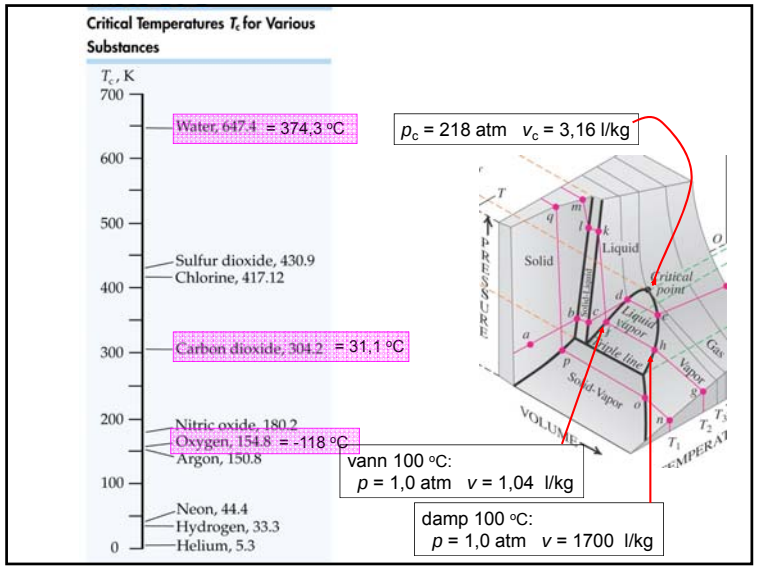


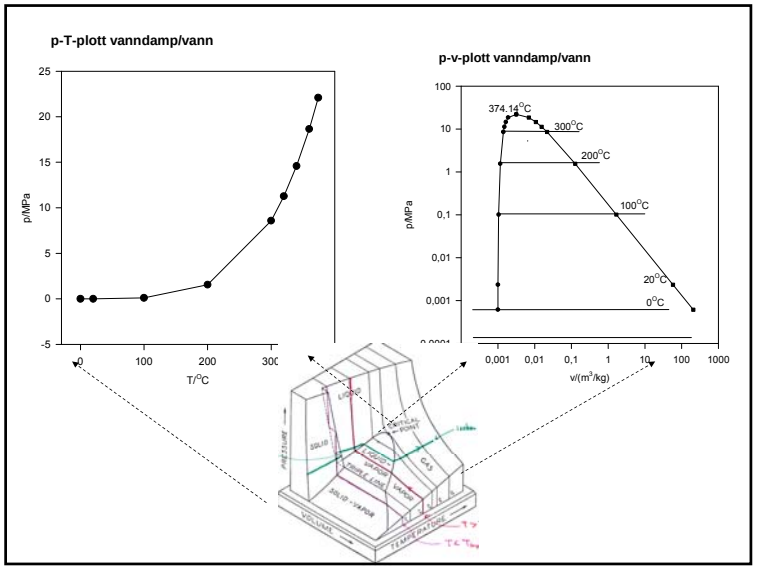
”Håndkokt” vann i sprøyte.



Fasediagram i *pT*-projeksjon





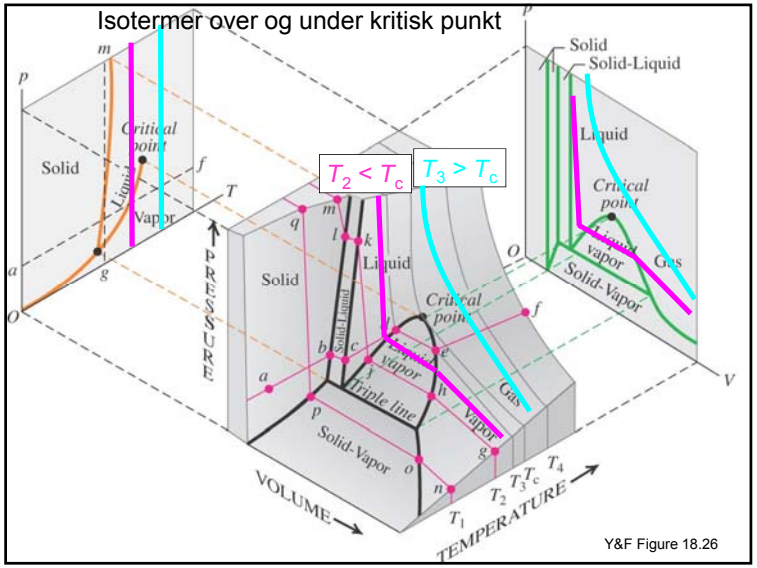


Data vanndamp og vann 0°C til 374°C

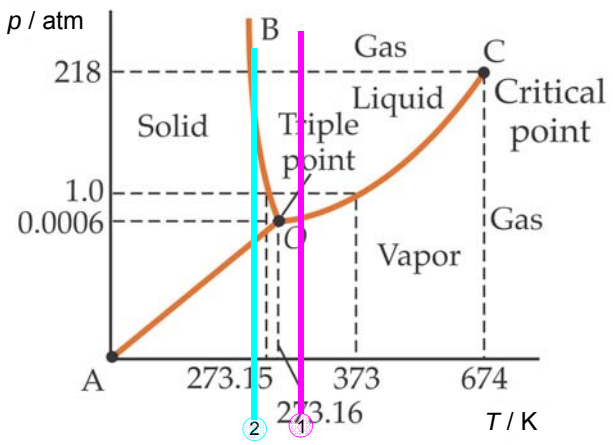
T/°C	p/MPa	v_v/(m³/kg)	v_g/(m³/kg)	L_f/(kJ/kg)	"R"
0,01	0,0006	0,001000	206,1	2501	8,31
20	0,0023	0,001002	57,8	2454	8,31
50	0,0123	0,001012	12,0	2383	8,23
100	0,1013	0,001044	1,67	2257	8,18
200	1,55	0,001157	0,127	1941	7,53
300	8,58	0,001404	0,0217	1405	5,84
320	11,27	0,001499	0,0155	1239	5,30
340	14,58	0,001638	0,0108	1028	4,62
360	18,65	0,001893	0,00695	721,0	3,66
374,14	22,09	0,003155	0,003155	0,0	1,97

p-T-plott

p = vanndampens metningstrykk
v_v = vannets spesifikke volum
v_g = vanndampens spesifikke volum
l_f = spesifikk fordampningsvarme
"R" = $p\,v_g/T$ (= R = 8,31 J/(K mol) hvis ideell gass)

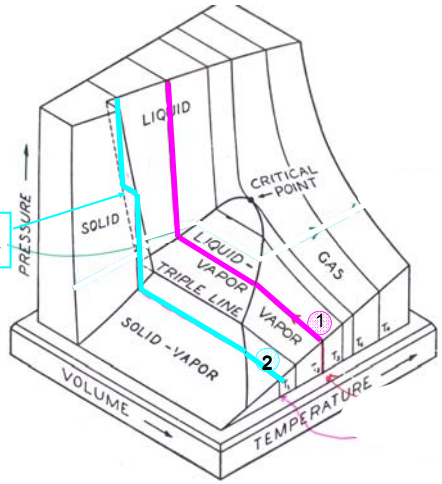


Fasediagram i *pT*-projeksjon **for H_2O**

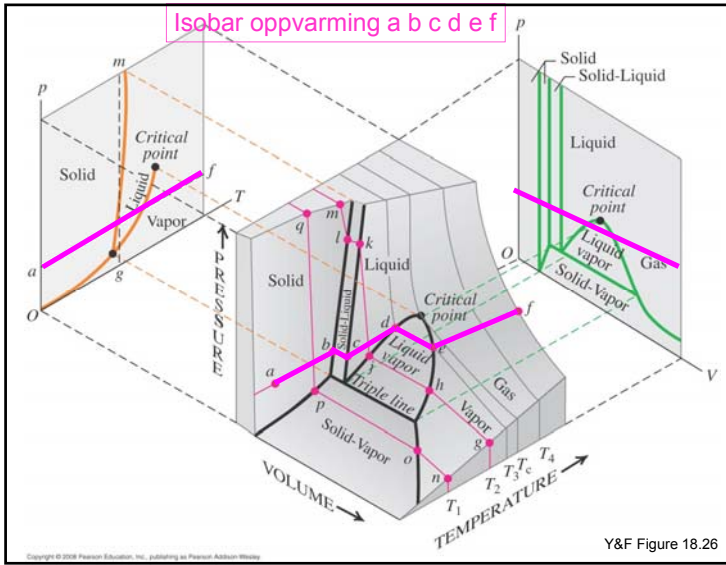


Fasediagram H_2O

is -> vann: volum avtar



Isobar oppvarming a b c d e f

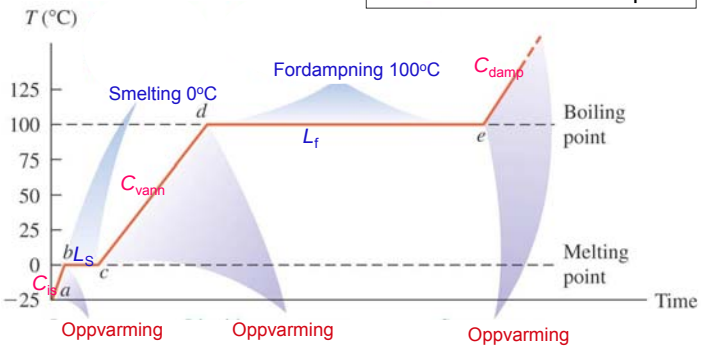


Y&F Figure 18.26

Isobar oppvarming (Smelting og fordamping)

Tilførsel av konstant varmeeffekt P til

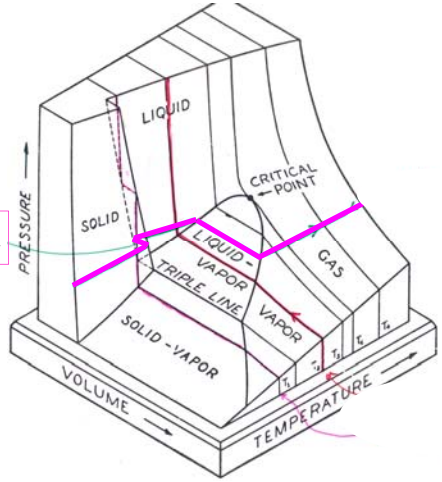
is → vann → vanndamp



Y&F Figure 17.21

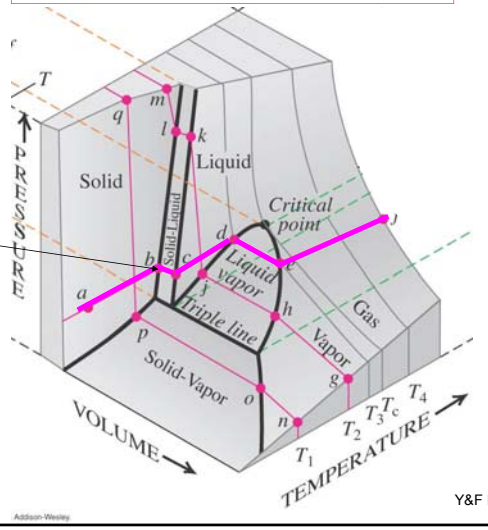
Isobar oppvarming VANN

is -> vann: volum avtar



Isobar oppvarming STOFFER ≠ VANN

volum øker



Y&F Figure 18.26

Faseoverganger. Oppsummering

- Fasediagram i *pVT*-rommet viser hvilke områder de tre faser fast, væske, gass kan eksistere hver for seg og sammen. Gjelder kun reine faser (én type stoff).
- I fasediagram i *pT*-projeksjon er sameksistensflatene kurver. Fasediagram i *pV*-projeksjon også ofte brukes.
- Smelting (fast -> væske): L'_s = spesifikk smeltevarme (J/kg)
- Fordamping (væske -> gass) : L'_f = spesifikk fordampingsvarme (J/kg)
- I *pT*-plott har sameksistenskurve væske/gass $dp/dT > 0$.
- I *pT*-plott har sameksistenskurve fast/væske $dp/dT > 0$, **unntatt H_2O** fordi is har større volum enn vann.
- Sameksistenskurve væske/gass har et maksimalt (kritisk) punkt (p_{krit}, T_{krit}). For $p > p_{krit}$ og/eller $T > T_{krit}$ har væske og gass samme egenskaper.